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Hydrogen Shoot a stream of photons through a container of hydrogen gas. Observe how photons of certain energies are absorbed, causing the electron to move to different orbits. Build the spectrum of hydrogen based on photons that are absorbed and emitted.

Bohr Model of Hydrogen Gizmo : Lesson Info :

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Shoot a stream of photons through a container of hydrogen gas. Observe how photons of certain energies are absorbed, causing the electron to move to different orbits. Build the spectrum of hydrogen based on photons that are absorbed and emitted.

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Gizmo Warm-up In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The Bohr Model: Introduction Gizmo™ explores this connection. The laser shown in the Gizmo can emit photons, or particles of light, at a variety of wavelengths.

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Student Exploration Bohr Model- Introduction (ANSWER KEY)

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The Bohr model of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the Bohr Model of Hydrogen Gizmo. The Gizmo shows a laser pointed at a container of

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hydrogen gas.

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Bohr model,
description of the
structure of atoms,
especially that of
hydrogen, proposed
(1913) by the Danish
physicist Niels Bohr.
The Bohr model of the
atom, a radical
departure from earlier,

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classical descriptions,
was the first that
incorporated quantum
theory and was the
predecessor of wholly
quantum-mechanical
models.

Bohr model | Description & Development | Britannica

Although revolutionary
at the time, the Bohr
model is a relatively
primitive model of the
hydrogen atom

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compared to the valence shell atom. As an initial hypothesis, it was derived as a first-order approximation to describe the hydrogen atom.

The Bohr Model | Introduction to Chemistry

2018 Gizmo Warm-up

In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related

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to the way electrons
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can emit photons, or
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The Bohr model is a
relatively primitive

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model of the hydrogen atom, compared to the valence shell atom model. As a theory, it can be derived as a first-order approximation of the hydrogen atom using the broader and much more accurate quantum mechanics and thus may be considered to be an obsolete scientific theory.

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